

## NOTE

**The Preparation of Beta-Chloropropionic Acid.**—Of all of the methods of preparation of  $\beta$ -chloropropionic acid described in the literature, only three appear suitable for the preparation of considerable quantities of the acid. Even these, however, have certain drawbacks, either giving poor yields or involving long tedious processes. The following is a convenient method for the preparation of this substance in good yield.

Fifty g. of trimethylene chlorohydrin is added drop by drop to 220 g. of concd. nitric acid contained in a narrow bottle immersed in water, the mixture being constantly stirred. The addition of the chlorohydrin requires approximately one hour. The solution is then warmed on the steam-bath for one hour, diluted with thrice its volume of water and extracted with four 60cc. portions of ether. The ether solution is dried with sodium sulfate and the ether distilled on the steam-bath. When most of the ether has been removed, a further quantity of nitrous fumes is evolved. The heating is then continued until no more liquid distills. The residue is distilled under diminished pressure, when a small quantity of oil and some nitric acid come over first. When no more brown fumes are observed, the temperature rapidly rises and the  $\beta$ -chloropropionic acid distills as a colorless, oily liquid; b. p., 127 (35 mm.). On standing, the acid solidifies to a mass of colorless crystals; yield, 32 g., or 56%.

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## ADDITIONS AND CORRECTIONS

1923, VOLUME 45

**Azido Carbondisulfide. I. Formation, Preparation and General Properties,** by A. W. Browne, A. B. Hoel, G. B. L. Smith and F. H. Swezey.

P. 2541. In line 7 of this article, instead of " $N_2-N-CS-S-S-CS-N-N_2$ ," read " $N_2=N-CS-S-S-CS-N=N_2$ ."

**Azido-dithiocarbonic Acid. I. Formation, Preparation and Properties,** by G. B. L. Smith and F. Wilcoxon with A. W. Browne.

P. 2605. In the last line of the text, instead of " $H-S-C(=S)-N=N=N$ ," read " $H-S-C(=S)-N=N=N$ ."

1924, VOLUME 46

**The Influence of Gelatin on the Transference Numbers of Hydrochloric Acid,** by Wesley G. France and Walter H. Moran.

P. 25. In Table IV, the values for  $C_1$  should read, 0.0814, 0.07475, 0.06990, 0.06473 and 0.05986; the values for  $C_2$  should read, 0.00924, 0.00199, 0.00021, 0.00008, and 0.00007.

P. 26. In Table V, the values for  $E_x$  should read, 0.00144, 0.00258, 0.00388 and